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IGCSE Chemistry Complete Revision Summary

Reactivity Series

- a) Atomic Structure
- b) Periodic Table
- c) Structure and Bonding
- d) Quantitative Chemistry
- e) Chemical Changes
- f) Energy Changes

- Reactivity of Metals
- Reactivity Series
- Extraction of Metals
- Acids and Bases
- Neutralization
- Making Soluble Salts
- Making Insoluble Salts
- Titrations
- Electrolysis
- Electrolysis of molten compounds
- Electrolysis of aqueous solutions
- Electrolysis of Aluminium

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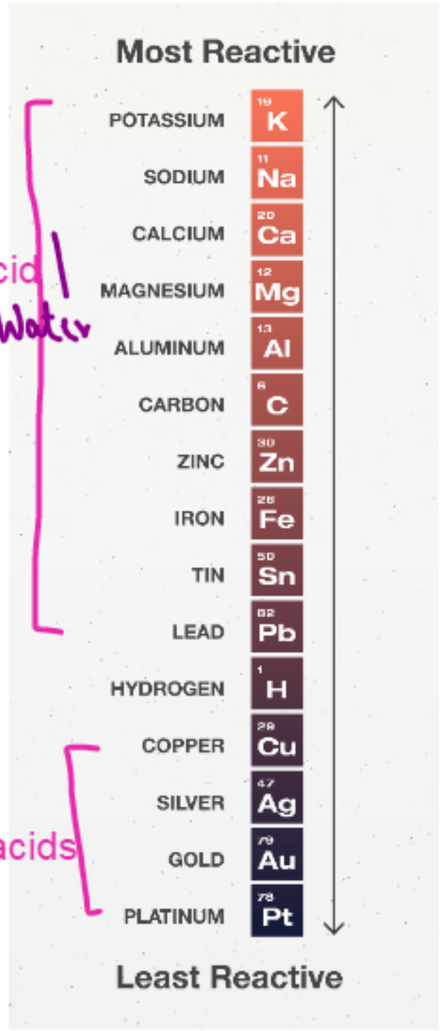
REACTIVITY SERIES

Metal + Dilute Acids = Salt + Hydrogen



React with Acid and displace Hydrogen

Water

Does not react with acids

or water

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DISPLACEMENT REACTION

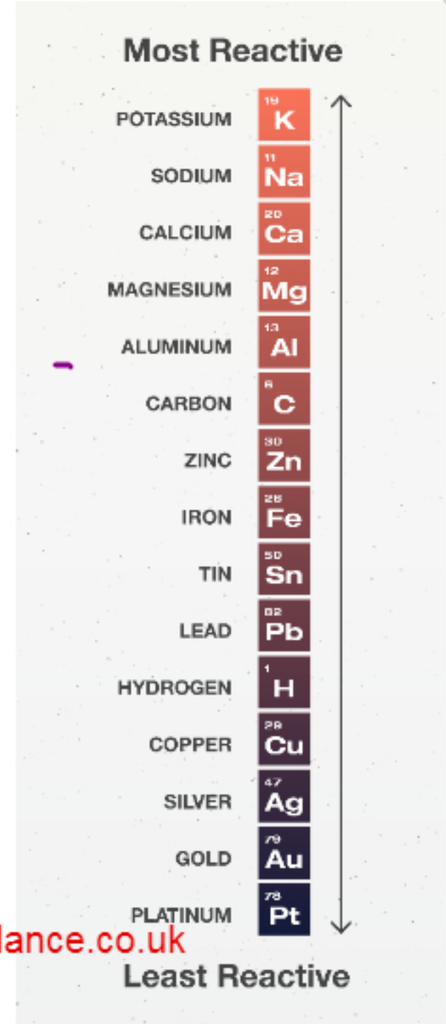
More reactive metal will displace the less reactive metal from its salt solution.

Magnesium + Zinc sulphate = Magnesium sulphate + Zinc

(More reactive metal) (Less reactive salt) (More reactive metal displaces the less reactive)

Lead + Magnesium Sulphate = No reaction

(Less reactive metal cannot displace the more reactive metal)



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KEY TERMS

Acids

Bases

Metals

Reactivity Series

Reduction

Oxidation

Alkali

Salt

Neutralization Reaction

Indicators

TEST YOURSELF

pH scale

Soluble Salts

Insoluble Salts

Electrolysis

Electrode

Anode

Cathode

Electrolyte

Ionic compounds

Cryolite

Bauxite

Ore

Metal Extraction

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